



# MASTER TUTORIALS

## Topic Name: Metals and NonMetals

### Assignment 1:

No Of Questions:20

1. Which of the following properties is not a characteristic of metals ?
  - (a) Metallic lusture
  - (b) High density
  - (c) Hardness
  - (d) Low melting and boiling point
2. Which of the following metals generally occur in liquid state ?
  - (a) Mercury
  - (b) Bromine
  - (c) Gallium
  - (d) A & C both
3. Reactivity of zinc is \_\_\_\_\_ than hydrogen.
  - (a) Less
  - (b) More
  - (c) Equal
  - (d) Sometimes more sometimes less
4.  $Zn + xHCl \longrightarrow ZnCl_2 + A$ , In above equation A & x are
  - (a)  $H_2$ , 2
  - (b)  $Cl_2$ , 1
  - (c)  $H_2$ , 3
  - (d)  $H_2$ , 4
5. When sodium reacts with cold water, then the product formed will be-
  - (a)  $Na_2O$
  - (b)  $NaOH$
  - (c)  $Na_2CO_3$
  - (d) All of these
6. What is the decreasing order of reactivity of following metals ?  
Na, Al, K, Cu, Ag, Fe
  - (a)  $Na > K > Al > Cu > Ag > Fe$
  - (b)  $K > Na > Al > Cu > Fe > Ag$
  - (c)  $K > Na > Al > Fe > Cu > Ag$
  - (d)  $K > Na > Al > Fe > Ag > Cu$
7. When a metal is added to dilute HCl solution, there is no evolution of gas. Metals is -
  - (a) K
  - (b) Na
  - (c) Ag
  - (d) Zn
8. On addition of which metal, copper sulphate solution (Blue colour) will be changed to colourless solution >
  - (a) Fe
  - (b) Ag
  - (c) Zn
9.  $Zn + H_2O$  (Steam)  $\longrightarrow A + B$   
In the above equation (a) and (b) are
  - (a) Zn &  $H_2$
  - (b)  $ZnH_2$  &  $O_2$
  - (c)  $ZnO_2$  &  $O_2$
  - (d)  $ZnO$  &  $H_2$
10. Which of the following metals reacts vigorously with oxygen?
  - (a) Zinc
  - (b) Magnesium
  - (c) Sodium
  - (d) Copper
11. Octet rule was given by -
  - (a) Rutherford
  - (b) Soddy
  - (c) Lewis & Kossel
  - (d) None of these
12. Exception of octet rule is -
  - (a) K
  - (b) Ca
  - (c) N
  - (d) He
13. Ionic bond is formed by -
  - (a) Loss of electrons only
  - (b) Gain of electrons only.
  - (c) Loss and gain of electrons both.
  - (d) Sharing of electrons.
14. Ionic bond is formed between -
  - (a) Two electropositive elements.
  - (b) Two electronegative elements.
  - (c) Electropositive & electronegative elements.
  - (d) None of these
15. During formation of ionic bond -
  - (a) There is force of repulsion between two negative ions.
  - (b) There is force of repulsion between two positive ions.
  - (c) There is force of attraction between positive & negative ions.
  - (d) None of these.
16. In the formation of ionic bond, cation is formed by-
  - (a) Gain of electron (s).
  - (b) Loss of electron(s).
  - (c) Sharing of electron(s)
  - (d) None of these
17. Ionic compound have –
  - (a) Low melting and high boiling points.
  - (b) High melting and low boiling points.

- (c) Low melting and low boiling points.
- (d) High melting and high boiling points.

**18.** Ionic compounds conduct electricity in-

- (a) Solid state
- (b) Fused state.
- (c) Gaseous state
- (d) Do not conduct electricity at all.

**19.** Ionic compounds are soluble in-

- (a) Water
- (b) Benzene
- (c) Ether
- (d) Alcohol

**20.** Physical nature of most of the ionic compounds is-

- (a) Solid
- (b) Liquid
- (c) Gas
- (d) May exist in any state.