



MASTER TUTORIALS

Topic Name: Metals and NonMetals

Assignment 1:

No Of Questions:20

- Which of the following properties is not a characteristic of metals ?
(a) Metallic lusture
(b) High density
(c) Hardness
(d) Low melting and boiling point
- Which of the following metals generally occur in liquid state ?
(a) Mercury
(b) Bromine
(c) Gallium
(d) A & C both
- Reactivity of zinc is _____ than hydrogen.
(a) Less
(b) More
(c) Equal
(d) Sometimes more sometimes less
- $\text{Zn} + x\text{HCl} \longrightarrow \text{ZnCl}_2 + \text{A}$, In above equation A & x are
(a) H_2 , 2
(b) Cl_2 , 1
(c) H_2 , 3
(d) H_2 , 4
- When sodium reacts with cold water, then the product formed will be-
(a) Na_2O
(b) NaOH
(c) Na_2CO_3
(d) All of these
- What is the decreasing order of reactivity of following metals ?
Na, Al, K, Cu, Ag, Fe
(a) $\text{Na} > \text{K} > \text{Al} > \text{Cu} > \text{Ag} > \text{Fe}$
(b) $\text{K} > \text{Na} > \text{Al} > \text{Cu} > \text{Fe} > \text{Ag}$
(c) $\text{K} > \text{Na} > \text{Al} > \text{Fe} > \text{Cu} > \text{Ag}$
(d) $\text{K} > \text{Na} > \text{Al} > \text{Fe} > \text{Ag} > \text{Cu}$
- When a metal is added to dilute HCl solution, there is no evolution of gas. Metals is -
(a) K
(b) Na
(c) Ag
(d) Zn
- On addition of which metal, copper sulphate solution (Blue colour) will be changed to colourless solution >
(a) Fe
(b) Ag
(c) Zn
(d) Hg
- $\text{Zn} + \text{H}_2\text{O (Steam)} \longrightarrow \text{A} + \text{B}$
In the above equation (a) and (b) are
(a) Zn & H_2
(b) ZnH_2 & O_2
(c) ZnO_2 & O_2
(d) ZnO & H_2
- Which of the following metals reacts vigorously with oxygen?
(a) Zinc
(b) Magnesium
(c) Sodium
(d) Copper
- Octet rule was given by -
(a) Rutherford
(b) Soddy
(c) Lewis & Kossel
(d) None of these
- Exception of octet rule is -
(a) K
(b) Ca
(c) N
(d) He
- Ionic bond is formed by -
(a) Loss of electrons only
(b) Gain of electrons only.
(c) Loss and gain of electrons both.
(d) Sharing of electrons.
- Ionic bond is formed between -
(a) Two electropositive elements.
(b) Two electronegative elements.
(c) Electropositive & electronegative elements.
(d) None of these
- During formation of ionic bond -
(a) There is force of repulsion between two negative ions.
(b) There is force of repulsion between two positive ions.
(c) There is force of attraction between positive & negative ions.
(d) None of these.
- In the formation of ionic bond, cation is formed by-
(a) Gain of electron (s).
(b) Loss of electron(s).
(c) Sharing of electron(s)
(d) None of these
- Ionic compound have -
(a) Low melting and high boiling points.
(b) High melting and low boiling points.

- (c) Low melting and low boiling points.
- (d) High melting and high boiling points.
- 18.** Ionic compounds conduct electricity in-
 - (a) Solid state
 - (b) Fused state.
 - (c) Gaseous state
 - (d) Do not conduct electricity at all.
- 19.** Ionic compounds are soluble in-
 - (a) Water
 - (b) Benzene
 - (c) Ether
 - (d) Alcohol
- 20.** Physical nature of most of the ionic compounds is-
 - (a) Solid
 - (b) Liquid
 - (c) Gas
 - (d) May exist in any state.